

Subatomic Particles

The three subatomic particles of the atom are the _____, _____, and _____. At the center of each atom lies the atomic _____ which consists of _____ and _____. The atomic number refers to the number of _____ in the nucleus. All atoms of the same element have the same number of _____, hence the atomic number. Isotopes are atoms that have the same number of _____ but a different number of _____. An isotope is identified by its atomic mass number, which is the total number of _____ and _____ in the nucleus. The number of neutrons is equal to the _____ minus the _____. A carbon isotope that has 6 _____ and 6 neutrons is identified as carbon-12, where 12 is the atomic mass number. A carbon isotope having 6 _____ and 8 _____, on the other hand is carbon-14.

Element Name	Symbol	Atomic Number	Mass Number	# of Protons	# of Electrons	# of Neutrons
Cobalt-59						
Chlorine-36						
Nitrogen-14						
Potassium-40						
Arsenic-75						
Gold-197						
Platinum-188						

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